

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

## Ionic & Covalent Compound Naming Race

First, identify whether these compounds are ionic or covalent. Then, use the correct formula writing rules to write the correct chemical formulas for each compound.

	Compound Name	Type of Compound: Ionic or Covalent	Chemical Formula
1)	copper (II) chlorite		
2)	sodium hydroxide		
3)	nitrogen dioxide		
4)	cobalt (III) oxalate		
5)	ammonium sulfide		
6)	aluminum cyanide		
7)	carbon disulfide		
8)	tetraphosphorous pentoxide		
9)	potassium permanganate		
10)	manganese (III) chloride		
	Compound Name	Type of Compound: Ionic or Covalent	Chemical Formula
11)	calcium bromate		
12)	carbon monoxide		
13)	potassium oxide		
14)	antimony tribromide		
15)	zinc phosphate		
16)	copper (II) bicarbonate		
17)	dinitrogen tetroxide		
18)	manganese (IV) carbonate		
19)	lead (IV) nitride		
20)	pentacarbon decahydride		

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## Ionic & Covalent Compound Naming

First, identify whether these compounds are ionic or covalent. Then, use the correct naming rules to write the correct names for each compound.

	Chemical Formula	Type of Compound: Ionic or Covalent	Compound Name
21)	<b>CdBr<sub>2</sub></b>		
22)	<b>Cr(Cr<sub>2</sub>O<sub>7</sub>)<sub>3</sub></b>		
23)	<b>SBr<sub>2</sub></b>		
24)	<b>(NH<sub>4</sub>)<sub>2</sub>CrO<sub>4</sub></b>		
25)	<b>CuO</b>		
26)	<b>Pt<sub>3</sub>(PO<sub>3</sub>)<sub>4</sub></b>		
27)	<b>Al(ClO<sub>4</sub>)<sub>3</sub></b>		
28)	<b>NH<sub>3</sub></b>		
29)	<b>Ca(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b>		
30)	<b>N<sub>2</sub>O</b>		
	Chemical Formula	Type of Compound: Ionic or Covalent	Compound Name
31)	<b>V(SO<sub>4</sub>)<sub>2</sub></b>		
32)	<b>Ag<sub>2</sub>CO<sub>3</sub></b>		
33)	<b>N<sub>2</sub>S<sub>3</sub></b>		
34)	<b>FeSO<sub>3</sub></b>		
35)	<b>Zn(NO<sub>2</sub>)<sub>2</sub></b>		
36)	<b>C<sub>6</sub>H<sub>12</sub>O<sub>6</sub></b>		
37)	<b>PCl<sub>3</sub></b>		
38)	<b>Mn(OH)<sub>7</sub></b>		
39)	<b>Ni(NO<sub>3</sub>)<sub>2</sub></b>		
40)	<b>O<sub>2</sub></b>		